

Module Handbook

Module Name:	Coordination Compounds
Module Level:	Bachelor
Abbreviation, if applicable:	KII 206
Sub-heading, if applicable:	-
Courses included in the module, if applicable:	-
Semester/term:	4 th / Second Year
Module coordinator(s):	HarsasiSetyawati, S.Si., M.Si
Lecturer(s):	Ahmadi Jaya Permana, S.Si., M.Si
Language:	BahasaIndonesia
Classification within the curriculum	Compulsory Course / Elective Studies
Teaching format / class hours per week during semester:	2 hours lectures (50 min / hour)
Workload:	2 hours lectures, 2 hour structural activities, 2 hours individual study, 13 week per semester, and total 78hours per semester ~ 2.6 ECTS *
Credit Points:	2
Requirements:	Inorganic Chemistry II
Learning goals/competencies:	<p>General Competence(knowledge) : Able to infer overall coordination compounds correctly</p> <p>Specific Competence :</p> <ol style="list-style-type: none"> 1. Able to explain the history and theory of the formation of coordination compounds coordination compounds 2. Capable of connecting formula or the chemical structure of the nomenclature of coordination compounds 3. Being able to calculate the effective atomic number of coordination compounds 4. Able to determine isomerization, thermodynamics and kinetics, and magnetic properties of coordination compounds 5. Being able to decipher Term Symbol and application of coordination compounds 6. Able to interpret spectra of coordination compounds.
Content:	In the course of coordination chemistry will be discussed about the history, nomenclature, effective atomic number, Term Symbol, spectra, isomerization, formation theory, thermodynamics and kinetics, magnetic properties, and applications of coordination compounds.
Attribut soft skill	Communication and Team-work skills
Study/exam achievements:	<p>Students are considered to be competent and pass if at least get 55 The final value is calculated as follows: Middle Semester Exam (UTS) = 35%; Final Semester Exam (UAS) = 35% Structured work = 10%; Presentation assignment (including assessing soft skills)= 20%</p> <p>Table Value Graduation A: 100> NA≥75 AB: 74.9> NA≥70</p>

	B: 69.9 > NA ≥ 65 BC: 64.9 > NA ≥ 60 C: 59.9 > NA ≥ 55 D: 54.9 > NA ≥ 40 E: 39.9 ≥ NA
Forms of Media:	Slides and LCD projectors, whiteboards, Internet
Learning Methods	<ul style="list-style-type: none"> - Lecture - Discussion - The assignment of creating a journal summary English with specific topics - Presentation in groups
Literature:	<ol style="list-style-type: none"> 1. Miessler, G.L., Tarr, D.A., 2014, Inorganic Chemistry, Prentice hall, Internanal Inc., New Jersey. 2. Jones, C., Thornback John., 2007, Medicinal Applications of Coordination Chemistry, RSC Publishing, United Kingdom. 3. Huheey, J. E., 1993, Inorganic Chemistry, Principles of structure and reactivity, 3rd ed., Harper dan Publisher, New York. 4. Basolo, Johnson, R., 1964, Coordination Chemistry, W.A. Benyamin, Inc., New York. 5. Burger, K., Coordination Chemistry : Experimental Methods, CRC Press, USA. 6. Davies, JA., Hockensmith, CM., Kukushkin, V Yu., Kukushkin. Yu N., 1996, Synthetic Coordination Chemistry, World Scientific, Singapore. 7. Figgis, Brian N., Hitchman, M A., 2000, Ligand Field Theory and Its Applications, Wiley-VCH, Canada. 8. The latest English Journal.
Notes:	*Total ECTS = {(total hours workload x 50 min) / 60 min } / 25 hours Each ECTS is equals with 25 hours