

Module Handbook

Module Name:	Inorganic Chemistry I (Practical)
Module Level:	Bachelor
Abbreviation, if applicable:	KII 202
Sub-heading, if applicable:	-
Courses included in the module, if applicable:	-
Semester/term:	3 rd / Second Year
Module coordinator(s):	Dr. Hartati, M.Si
Lecturer(s):	Harsasi Setyawati, S.Si., M.Si; Ahmadi Jaya P, S.Si., M.Si
Language:	Bahasa Indonesia
Classification within the curriculum	Compulsory Course / Elective Studies
Teaching format / class hours per week during semester:	2 hours laboratory work (50 min / hours)
Workload:	2 hours doing worksheet and pretest preparation, 2 hours laboratory work, 2 hours group discussion, searching literature and writing report, 13 week per semester, and total 78 hours per semester ~ 2.6 ECTS *
Credit Points:	1 SCU
Requirements:	Inorganic Chemistry I
Learning goals/competencies:	<p>General Competence (skill) : Being able to do the synthesis, purification, and qualitative analysis of inorganic salts non-metallic elements with attention to hygiene and safety in the laboratory as well as being able to identify a few examples of the molecular structure of inorganic salts non-metallic elements</p> <p>Specific Competence :</p> <ol style="list-style-type: none"> 1. Able to perform the synthesis, purification and analysis of qualitative compounds of hydrogen, boron, nitrogen, iodine, carbon, chlorine, bromine, sulfur 2. Be able to identify a few examples of the molecular structure of inorganic salts non-metallic elements using hybridization theory.
Content:	In the course of Inorganic Chemistry Lab 1 will be discussed on structure, synthesis, purification and analysis of qualitative salts of single, double salt, the properties of halogen compounds, oxygen, sulfur, boron and nitrogen
Attribut soft skill	Discipline and team-work
Study/exam achievements:	<p>Students are considered to be competent and pass if at least get 55</p> <p>Final Exam Practice 40%; softskill 10%, 50% Daily practical value</p> <p>Table Value Graduation A: 100> NA≥75 AB: 74.9> NA≥70 B: 69.9> NA≥65</p>

	BC: 64.9 > NA ≥ 60 C: 59.9 > NA ≥ 55 D: 54.9 > NA ≥ 40 E: 39.9 ≥ NA
Learning Methods	Practical in Laboratory, discussion, Demonstration of the model structure
Form of Media	Laboratory equipments
Literature:	<ol style="list-style-type: none"> 1. Angelici, R.J., 1986, <i>Synthesis and technique in Inorganic in Inorganic Chemistry</i>, University Science Books, California. 2. Pass, G., and Sutcliffe, H., 1974, <i>Practical Inorganic Chemistry</i>, 2nd ed., Chapman and Hall, London. 3. Szafran, Z., Pike, M. R., 1991, <i>Miroscale Inorganic Chemistry, a Comprehensive Laboratory experience</i>, John Willey and Sons.
Notes:	*Total ECTS = {(total hours workload x 50 min) / 60 min } / 25 hours Each ECTS is equals with 25 hours